

Comment

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Comment on "Free-radical formation by the peroxidase-like catalytic activity of  $MFe_2O_4$  (M = Fe, Ni and Mn) nanoparticules"

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The peroxidase-like activity of magnetite and its connection with a neurodegenerative disease such as Alzheimer's disease has been recently questioned by our group,<sup>1</sup> despite claims by Gao *et al.*<sup>2</sup> Since reproducibility is the essential factor in experimental sciences, the article of Moreno Maldonado *et al.* on the peroxidase-like activity of magnetite attracted our attention.<sup>3</sup> As a matter of fact this new report is providing additional evidences on the absence of peroxidase activity of magnetite derivatives. The following comments are not a question of semantics, but are driven by the necessity to correctly

describe the different possible origins of the oxidative stress involved in many neurodegenerative pathologies.

*First question*: are the authors really observing a "peroxidase-like catalytic activity" as claimed in the title? Peroxidases are enzymes able to perform the 2-electron oxidation of a substrate AH<sub>2</sub> using hydrogen peroxide as oxidant (see equation 1).

$$AH_2 \xrightarrow{H_2O_2} A + 2 H_2O$$
 (Equation 1)

. . .

Mechanistic details concerning peroxidases and biomimetic models are available in text-books.<sup>4,5</sup> The authors did not report the catalytic oxidation of a peroxidase substrate, but they detected by spin-trap methods the presence of two radicals, HO• or HOO•, derived from the reaction of hydrogen peroxide or molecular oxygen, respectively, with low-valent metal ions. The hydroxyl radical HO• is the one-electron reduction product of hydrogen peroxide, generated by a low-valent metal ion in a Fenton-type reaction (see equation 2).

 $H_2O_2 \xrightarrow{M^{n+}} HO^{\bullet} + HO^{-} \quad (Equation 2)$ 

The hydroperoxyl radical HOO• is the protonated form of the superoxide anion  $O_2 \bullet^-$ , which is the result of the one-electron reduction of molecular oxygen by a low-valent metal ion (equation 3). The protonation of superoxide anion with a pKa value of 4.9 generates the neutral radical HOO• (named hydroperoxyl or perhydroxyl).<sup>6</sup>

$$O_2 \xrightarrow{\mathsf{M}^{\mathsf{n}+}} O_2^{-\bullet} \xrightarrow{\mathsf{H}^+} \mathsf{HOO}^{\bullet} \quad (Equation 3)$$

These two oxygen-centered radicals are clearly not the products of a peroxidase-type reaction.

Second question: are both oxygen-centered radicals generated by a catalytic reaction? Peroxidases are highly efficient catalysts. For instance, a  $2\times10^{-11}$  M concentration of the paradigm horseradish peroxidase (HRP) achieves oxidation of  $4\times10^{-5}$  M of tetramethylbenzidine within 5 min, in acetate buffer pH 3.5, thus giving rise to a turnover number of 2 millions of catalytic cycles within this short period of time, which corresponds to a turnover rate of 7000 cycles per second.<sup>1</sup> In the article by Moreno Maldonado *et al.*, the total production of HO• + HOO• with magnetite was slightly below 2000 nM (lanes "Fe-MNP", Figure 4 of the ref. 3) for an initial concentration of Fe(II) of 1600  $\mu$ M (120  $\mu$ g of Fe<sup>II</sup><sub>0.81</sub>Fe<sup>III</sup><sub>2.19</sub>O<sub>4</sub> in 260  $\mu$ L of solution, Experimental

Methods). So the yield of these two oxyl radicals was only 0.13% with respect to Fe(II), a sub-stoechiometric number, far from a catalytic activity. The reported initial concentration of  $H_2O_2$  was 0.49 M (Experimental Methods). The production of 2  $\mu$ M of HO• + HOO• therefore corresponds to the conversion of 0.0004%, a very small fraction of the starting amount of  $H_2O_2$ . So, the very low and non-catalytic yield of HO• + HOO• with respect to Fe(II)] is probably the result of a Fenton-type chemistry induced by adventitious redox-active iron impurities contained in the isolated nanoparticles or generated by traces of reductants.

In conclusion, the experimental results reported by Moreno Maldonado *et al.*, provide an additional proof that magnetite nanoparticles has no peroxidase activity. In addition, the authors mentioned that "no radical formation was evidenced for any sample at pH = 7.4", a value corresponding to physiological pH. So, the biomedical claims are not consistent with the reported results. Consequently, the title of the article and all comments concerning a "peroxidase-like catalytic activity" (and putative biomedical claims) of magnetite and other ferrite nanoparticles, did not resist to a careful examination of experimental results. All the reported results can be summarized by the following sentence: "*No peroxidase activity is observed with*  $MFe_2O_4$  (M = Fe, *Ni and* Mn) *nanoparticules and only traces of hydroxyl and hydroperoxyl radicals are observed at* pH = 4.8. At physiological pH, these radicals are not detected".

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